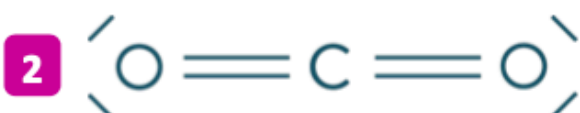


## Correction Activité documentaire n°9.1 : Schéma de Lewis

**1** Azote :  $1s^2 2s^2 2p^3$     soufre :  $1s^2 2s^2 2p^6 3s^2 3p^4$



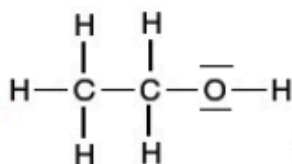
Ce schéma est correct car les atomes de carbone et d'oxygène doivent être entourés de 4 doublets liants et non-liants.

**3**

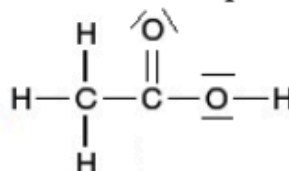
Molécule	Formule	Schéma de Lewis
Eau	H <sub>2</sub> O	$\text{H}-\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{O}}}-\text{H}$
Dioxygène	O <sub>2</sub>	$\langle \text{O}=\text{O} \rangle$
Chlorure d'hydrogène	HCl	$ \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}}-\text{H}$

**4**

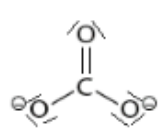
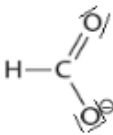
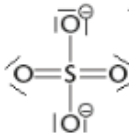
**Ethanol**



**Acide éthanoïque**



**5**

Ion carbonate CO <sub>3</sub> <sup>2-</sup>	Ion hypochlorite ClO <sup>-</sup>	Ion méthanoate HCOO <sup>-</sup>	Ion sulfate SO <sub>4</sub> <sup>2-</sup>	Ion ammonium NH <sub>4</sub> <sup>+</sup>	Ion oxonium H <sub>3</sub> O <sup>+</sup>
	$ \overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{Cl}}}-\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{O}}}$			$\text{H}-\overset{\oplus}{\text{N}}(\text{H})-\text{H}$	$\text{H}-\overset{\oplus}{\text{O}}(\text{H})-\text{H}$